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Chapter 21: ELECTROCHEMISTRY TYING IT ALL TOGETHER-

RT In  $K = G = -nFE_o$  CH 17-20 = CH15 = CH 21. 2 Important errata

to fix.  $F = 9.649 \times 10^4$  C/mole in problem on current measurement.

Beginning of section on anodes and cathodes should read: " Oh, and by

Chapter 21: ELECTROCHEMISTRY TYING IT ALL TOGETHER

Chapter 21 = Electrochemistry 21.1 = Voltaic Cells (Day 12 = A-4/24,

B-4/27) 21.2 = Types of Batteries (Day 12 = A-4/24, B-4/27) 21.3 =

Electrolysis (Day 13 = A-4/28, B-4/29) Chapter 21 Homework (Due

Day 14 = A-4/30, B-5/1) 21.1 - 21.2 = #26-#36, #43-#46 21.3 =

#20-#24 Unit #14 Test = Redox & Electrochemistry (Day 14 = A-4/30, B-5/1)

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A.P. Chemistry Practice Test - Ch. 17: Electrochemistry MULTIPLE

CHOICE. Choose the one alternative that best completes the

statement or answers the question. 1) The gain of electrons by an

element is called \_\_\_\_\_. A)oxidation B)sublimation C)reduction

D)disproportionation E)fractionation 2) \_\_\_\_\_ is reduced in the

following reaction:

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### Topic 13 - Electrochemistry - A-Level Chemistry

Chapter 21 Electrochemistry Section 21.1. Electrochemistry represents  
the interconversion of chemical energy and electrical energy.  
Electrochemistry involves redox (reduction-oxidation) reactions  
because the electrical energy (flow of electrons) has at its origin the  
oxidation (loss of electrons) and reduction (gain of electrons) of  
species.

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Electrochemistry class 12 important questions | Chemistry

Redox in Electrochemistry Electrochemistry is the study of the redox processes by which chemical energy is converted to electrical energy and vice versa. Electrochemical processes are useful in industry and critically important for biological functioning. In Chapter 19, you read that all redox reactions involve a transfer of

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