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Pearson Chemistry © 2012 : Program Components

1 a) The relative atomic mass of an element is the weighted average mass of the isotopes of the element.
(1) It is measured on a scale on which a carbon-12 atom has a mass of exactly 12 units. (1). b) The total mass of 100 atoms = $(75 \times 35) + (25 \times 37) = 3550$ (1) The average mass of 1 atom = $3550/100 = 35.5$
(1)

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